



## Rainfall chemistry: long range transport versus below cloud scavenging. A two-year study at an inland station (Opme, France)

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Guillaume Bertrand, Hélène Celle-Jeanton, Paolo Laj, J. Rangognio, Gilles Chazot. Rainfall chemistry: long range transport versus below cloud scavenging. A two-year study at an inland station (Opme, France). *Journal of Atmospheric Chemistry*, 2009, 60 (3), pp.253-271. 10.1007/s10874-009-9120-y . insu-00418691

**HAL Id: insu-00418691**

**<https://hal-insu.archives-ouvertes.fr/insu-00418691>**

Submitted on 26 Jun 2012

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## 35 1. Introduction

The chemical composition of precipitation is strongly linked with aerosols (both particles and gases) carried by rainwater (Junge, 1963). They are removed from the atmosphere by two scavenging processes (Meszaros, 1981; Huang et al., 2008): 1) the rainout corresponds to  
40 condensation processes within the clouds during their formation and travel, 2) the washout is characterized by the impaction with falling raindrops. In this way, the rainfall signature integrates both the long-range transport of chemical species by clouds and the local scavenging of atmospheric aerosols during the rain event. These figures vary strongly both  
45 from event to event and within a single event because rainout and washout proportions strongly depend on the environment of the sampling site (urbanized, rural or remote). It is a function of the aerosol loading and its vertical distribution in the lower atmosphere (Lim et al., 1991). Therefore the composition of rainwater depends both on the meteorological context, which influences clouds and rains characteristics and on local/long-range transport of aerosols and their spatial distribution (Durana et al., 1992). The sources of dissolved  
50 components in rainwater has often been classified in different categories based on air mass origin (Celle, 2000): 1) marine source which essentially provides  $\text{Na}^+$  and  $\text{Cl}^-$ , 2) terrestrial source e.g. soil dust ( $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{HCO}_3^-$ ), 3) anthropogenic sources e.g. traffic, industry, agriculture and burning of vegetation mainly associated with  $\text{NO}_3^-$ ,  $\text{K}^+$ ,  $\text{SO}_4^{2-}$ ,  $\text{NH}_4^+$ . Moreover the atmosphere constitutes a reactive medium. Some aerosols can interact to form  
55 new species in dry air before their dissolution in the meteoric water. This latter can also induces further reactions due to the presence of ionic species. These interactions result in a wide range of chemical composition, as well as in variations of pH (Berner and Berner, 1987; Cyrus et al., 1995).

Sources and chemical interactions are difficult to discriminate analytically but their relative  
60 contributions could be estimated by a coupling of chemical composition and meteorological

data (Beverland et al., 1998; Avila et Alarcon, 1999; Celle, 2000; Huang et al., 2008; Celle-Jeanton et al., 2009). In this purpose, 217 precipitation events and 204 infra-event samples have been collected during the November 2005-October 2007 period at Opme meteorological station (Massif Central, France) and were characterised as a function of their air mass back-trajectories.

The objectives of such an approach are: 1) to evaluate the role of the long-range transport and characterize the distant sources of aerosols and 2) to determine how local meteorological variables and environmental features could influence the chemical signature of rainwater at a small temporal scale.

## 2. Methodology

### Sampling site

The sampling device has been settled at Opme meteorological station (45°43'N, 3°5'30''E, z = 650 m) belonging to Observatoire de Physique du Globe de Clermont-Ferrand. The meteorological station is located 8 km south of Clermont-Ferrand urban district, characterised by a population of 260 000 inhabitants (INSEE, 1999), and which constitutes the main regional activity centre (industries, tertiary sector). Surrounding this area, land occupation mainly consists on cereal-growing in the eastern part (Limagne basin) and on rearing or forest on the western part (Chaîne des Puys). The lithology mainly consists of sedimentary deposits on the Limagne basin and crystalline basement (volcanic and granitic) on the west (Fig.1).

In addition to the collection of rainwater, the recorded parameters are: rainfall intensity (mm/h), temperature (°C), wind speed and direction. Daily rainfalls have been collected through an automatic precipitation collector (Eigenbrodt NSA 181/KHS) with a 500 cm<sup>2</sup>

aperture. A precipitation sensor causes the cover device to open up the collection funnel at the start of precipitation and then avoids dry depositions in wet samples. Precipitation flows from the funnel down a bottle tray built in with 9 collecting bottles of 1 litre each (one bottle for each day of the week, the additional 9<sup>th</sup> bottle is thought as an overflow for extremely high precipitation). When precipitation has ceased and after evaporation of the sensor surface, the funnel lids closed automatically. The system is refrigerated at 5°C and allows a good conservation of samples that are collected once a week. After the sampling, the bucket is systematically cleaned with distilled water. It has to be noted that some events were not analysed for they are too small in volume (analyses require a minimum rainfall volume of 15 ml). Moreover, due to the defection from the cover device, some events were not sampled. In spite of this, the sampling of 89% of the total rainy days of the period has been performed.

The sequential sampling of rainwater was carried out using a modified version of Bourrié's device (1978; Celle-Jeanton et al., 2004) which is located near the total rain collector. This collector of 1134 cm<sup>2</sup> area is feeding a set of 10 vials of 100 ml which are successively filled by the way of a three-way pipe. It enables the collection of rains up to 17.6 mm. Laboratory tests assessed that mixing between successive vials is about 3% (Rangognio, 2006).

#### Analytical methods

Upon sampling, the rainwater was taken to the laboratory where electrical conductivity (EC) and pH are measured with WTW Multi 340i. HCO<sub>3</sub><sup>-</sup> concentrations are determined by titration with a 5.10<sup>-4</sup> M solution of sulphuric acid. Samples were subsequently poured into standard 100 ml polyethylene flask for further analysis and stored at 4°C for further chemical analysis of anions (SO<sub>4</sub><sup>2-</sup>, Cl<sup>-</sup>, NO<sub>3</sub><sup>-</sup> and PO<sub>4</sub><sup>3-</sup>) and cations (Ca<sup>2+</sup>, Na<sup>+</sup>, NH<sub>4</sub><sup>+</sup>, K<sup>+</sup>, Mg<sup>2+</sup>). Concentrations of ionic species were determined by ion chromatography, using a DIONEX

DX320 chromatograph with a AS11 column for anions. This system uses an eluent generator to convert pure water into potassium hydroxide eluent through electrolysis for anions separations. A DIONEX ICS500 chromatograph with a CS16 column was used for cations determination. A methanesulfonic acid eluent was used for cation separations. The analysis method is similar than the one described by Jaffrezo et al. (1998) and Ricard et al. (2002). The flow rates of eluents were 1 ml/min. The detection limits are 0.05  $\mu\text{eq/l}$  for  $\text{Cl}^-$ , 0.1  $\mu\text{eq/l}$  for  $\text{NO}_3^-$ , 0.6  $\mu\text{eq/l}$  for  $\text{SO}_4^{2-}$  and  $\text{PO}_4^{3-}$ , 0.06  $\mu\text{eq/l}$  for  $\text{Na}^+$ , 0.04  $\mu\text{eq/l}$  for  $\text{NH}_4^+$ , 0.3  $\mu\text{eq/l}$  for  $\text{K}^+$ , 0.1  $\mu\text{eq/l}$  for  $\text{Mg}^{2+}$  and  $\text{Ca}^{2+}$ . The uncertainty of these chromatograph systems is 5%. In order to assess the validity of the sampling device, blanks were determined for the analysed elements. The highest value were 1.0  $\mu\text{eq/l}$  for sodium and 0.7  $\mu\text{eq/l}$  for calcium, the other elements were not detected. The charge balance between anions and cations was assessed through linear regression analyses of the data. The results yield value of 0.93 for the regression coefficient indicating the completeness of measured parameters.

### 3. Results and discussion

#### Rainfall main characteristics

Table 1 shows the statistics calculated for the chemical parameters measured in the rain samples collected at Opme from Nov-2005 to Nov-2007. Minimum and the maximum values of each chemical parameter emphasize the high variability of the chemical composition and mineralization of rainwater. This variability can be due to the rain amount (Al-Momani et al., 1995; Khwaja and Husain, 1990; Hicks and Shannon, 1979) as can be seen on Fig. 2 or to the influence of various chemical sources. In this context, one way to keep out the influence of rain quantity is to use the volume-weighted mean of concentrations to characterize the general features of the precipitation chemistry.

At Opme meteo-station, the volume-weighted mean pH is of 5.1, showing a tendency of slight acidity of rainwater. Generally, unpolluted precipitations are known to be weakly acid due to their equilibrium with atmospheric CO<sub>2</sub>. Then, the pH of chemically undisturbed rainwater is about 5.7 (Berner and Berner, 1987). The acidic feature is common in industrialised countries (Tost et al., 2007) and is due to anthropogenic emissions such as SO<sub>x</sub> and NO<sub>x</sub>. The volume weighted mean (VWM) concentrations of the major ionic species are in the following order: Ca<sup>2+</sup> > HCO<sub>3</sub><sup>-</sup> > NO<sub>3</sub><sup>-</sup> > SO<sub>4</sub><sup>2-</sup> > NH<sub>4</sub><sup>+</sup> > Na<sup>+</sup> > Cl<sup>-</sup> > H<sub>3</sub>O<sup>+</sup> > Mg<sup>2+</sup> > K<sup>+</sup> > PO<sub>4</sub><sup>3-</sup>. Rainfalls chemistry appears then to be mainly influenced by elements coming from terrestrial and anthropogenic sources. Except for SO<sub>4</sub><sup>2-</sup>, Negrel and Roy (1998) found similar concentrations at Sainte-Marguerite site (12 km to the east from Opme). However concentrations of SO<sub>4</sub><sup>2-</sup> analysed during the period Nov-2005 to Nov-2007 are two times lower than the ones of Sainte-Marguerite measured in the period Mar-1994 to Apr-1995. This difference could be explained by the general decrease in sulphur emissions between 1986 and 2001 observed in Europe (Fowler et al., 2005); in Great Britain, the sulphate concentrations in rain drops decrease of 47% during the 1986-2001 period; in continental Europe reductions in SO<sub>2</sub> emissions of decrease by 72% during the same period.

#### Source and processes affecting the rainwater mineralization

To evaluate the main geochemical sources of dissolved compounds in the rainwater, inter-species relationships could be investigated through a multivariate statistical analysis. A factorial analysis coupled with varimax rotation permits to determine the preferential association between the chemical species in rain (Plaisance et al., 1996; Simeonov et al., 2003; Zunckel et al., 2003). Principle factor analysis was applied to the 217 chemical analyses. Consecutive factors are uncorrelated or orthogonal to each other. Factor loadings correspond to the correlations of each variable with the factor. Therefore, the more the

absolute value approaches one, the more the variable is correlated with the factor. A negative  
variable loading indicates that it varies in opposite direction to the variables with positive  
signs. Three factors have been used in this analysis, giving each variable a loading within  
each factor. Loadings greater than 0.5 are considered to be significant components of the  
factor.

The three factors explain 99.7% of the total variance of all the data (Table 2). The first factor  
(41.7% of the total variance) suggests the association of anthropogenic species ( $\text{NO}_3^-$ ,  $\text{SO}_4^{2-}$ ,  
 $\text{NH}_4^+$ ) with  $\text{Ca}^{2+}$  mainly of a terrestrial origin. The relatively high loading in  $\text{Mg}^{2+}$  (0.480)  
allows associating this element to factor 1, that then represents neutralisation process. In fact,  
these well-correlations between ions result from atmospheric chemical reactions, probably  
from the reaction of the  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$  acids with alkaline compounds rich in  $\text{Ca}^{2+}$  and  
 $\text{Mg}^{2+}$  carried into the atmosphere. Neutralisation of  $\text{HNO}_3$ , and  $\text{H}_2\text{SO}_4$  by soil-derived  
particles and by ammonia emissions can form aerosols such as  $\text{NH}_4\text{HSO}_4$ ,  $(\text{NH}_4)_2\text{SO}_4$ ,  
 $\text{NH}_4\text{NO}_3$ ,  $\text{Ca}(\text{NO}_3)_2$  and  $\text{CaSO}_4$  (Sisterson, 1989; Hov and Hjollo, 1994). The second factor  
(35.6% of the total variance) highlights the association of  $\text{Na}^+$ ,  $\text{Cl}^-$  and  $\text{Mg}^{2+}$  and then  
corresponds to the marine origin. The third factor associates calcium and bicarbonates (22.4%  
of the total variance). This factor 3 expresses the control of the terrestrial source that could  
influence the alkalinity of the atmospheric water due to carbonate dissolution.

Many studies pointed out the role of meteorological factors in determining the chemical  
features of precipitations (Beverland et al., 1998; Avila and Alarcon, 1999; Celle-Jeanton et  
al., 2009). To investigate the origin of components, the chemical content of the 217  
precipitation events sampled during our study have been coupled with the corresponding air  
mass back trajectories.



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Air mass back trajectories were calculated for each sample by using the NOAA HYbrid Single-Particle Lagrangian Integrated Trajectory Version 4 (HYSPLIT4) model (Draxler and Rolph, 2003; Rolph, 2003). The back trajectories were calculated for a three days period and for three altitudes: 900 m, 1500 m and 3000 m. Dataset of winds' speed and direction (1981-  
 195 2000, MétéoFrance) allows calculating the main air masses origin and thus permits to approach the weather patterns of Clermont-Ferrand. Four main transport patterns have been segregated: sector 1 that consists of air masses coming mainly from Atlantic Ocean (22.5% of air-masses origin), sector 2 that corresponds to Northern and Central Europe (28.3%), sector 3 towards the South including Iberian Peninsula and Northern Africa (25.8%), a regional sector  
 200 4, with a corresponding wind's speed lower than 2 m/s, corresponds to 22.4% (Fig. 3). Chemical contents for each sector were determined taking into account the role of the different sources of mineralization highlighted by the factor analysis previously depicted. In particular, some ions result from the contribution of both sea-salts (ss) aerosols and non-sea-salt (nss) species. The sea-salt input is determined by assuming that all  $\text{Na}^+$  (Brewer, 1975) is  
 205 provided by the marine source and that the proportionate amount of ions as  $\text{Cl}^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{Mg}^{2+}$ ,  $\text{K}^+$ ,  $\text{Ca}^{2+}$ , are derived from sea salts.

Statistical results and distribution of pH are presented on Table 3 and Fig. 4, respectively. In order to determine the preferential neutralisation process, the linear regression between  $\text{nssCa}^{2+}$  (provided by terrestrial dusts such as  $\text{CaCO}_3$  or  $\text{CaMg}(\text{CO}_3)_2$ ) or  $\text{NH}_4^+$  with acidic  
 210 species ( $\text{NO}_3^-$  and  $\text{nssSO}_4^{2-}$ ) are reported on Fig. 5.

The chemical content of rainfalls coming from sector 1 is dominated by non sea-salt calcium ( $\text{nssCa}^{2+} = 23.9 \mu\text{eq/l}$ ) and bicarbonates ( $24.1 \mu\text{eq/l}$ ), but is also influenced by anthropogenic species ( $[\text{NH}_4^+] = 20 \mu\text{eq/l}$ ,  $[\text{nssSO}_4^{2-}] = 16.8 \mu\text{eq/l}$ ,  $[\text{NO}_3^-] = 16.0 \mu\text{eq/l}$ ). Sodium and

215 chloride concentrations show the influence air masses coming from the Atlantic Ocean. Whereas a major marine source could be expected featuring atlantic events, it comes out that terrestrial and anthropogenic sources mark their chemical feature. This observation is in good agreement with the study of Rahn et al. (1982), who show that aged maritime air masses often contain a high proportion of submicron crustal and anthropogenic particles due to their long  
220 atmospheric residence times. The high content of calcium in rainfall was also shown, for rain events on the Atlantic coast, by Beysens et al. (2006) who indicated that the source of high concentration in calcium could be due to the removal of soils particles but also originate from agriculture activities. It would be consistent with VWM ammonium concentrations (20.0  $\mu\text{eq/l}$ ). Although 4 samples are acidic ( $\text{pH} < 4.5$ ) and 9 have  $\text{pH} > 6.5$ , 31 events have a nearly  
225 neutral pH. Fig. 5 shows the good relationships between  $\text{nssCa}^{2+}$  and  $(\text{NO}_3^- + \text{nssSO}_4^{2-})$ ,  $\text{NH}_4^+$  and  $(\text{NO}_3^- + \text{nssSO}_4^{2-})$ ,  $R = 0.77$  and  $0.74$ , respectively, and attests consequently that neutralisation can frequently occur in air-masses coming from the sector 1.

Rainwater coming from the sector 2 presents high concentrations of  $\text{NO}_3^-$ ,  $\text{nssSO}_4^{2-}$  and  $\text{NH}_4^+$   
230 (respectively 52.3  $\mu\text{eq/l}$ , 35.6  $\mu\text{eq/l}$ , 33.3  $\mu\text{eq/l}$ , Table 3) and the highest total mineralization of the four sectors. This group also exhibits the highest VWM in non sea-salt calcium (53.2  $\mu\text{eq/l}$ ) due probably to the removal of Ca-rich dusts. Such influence was highlighted by Sanusi et al. (1996) in northern France; they particularly pointed out the role of loess that can be easily removed from the Alsatian basin. High mineralization correlated to high content in  
235 anthropogenic components let appear that condensation nuclei are numerous and mainly constituted of polluted aerosols. An increase in cloud condensation nuclei concentration causes an increase in the number of cloud droplets (Reade et al., 2006).

Low mean pH of 4.8, (Table 3, Fig. 4) indicates potential acid influences of air-masses coming from this sector. This feature is consistent with the low value of correlation

240 coefficient ( $R = 0.46$ , Fig. 5) between  $\text{nssCa}^{2+}$  and  $[\text{NO}_3^- + \text{nssSO}_4^{2-}]$ . Concerning the interaction between  $\text{NH}_4^+$  and  $[\text{NO}_3^- + \text{nssSO}_4^{2-}]$ , the correlation is better ( $R = 0.73$ ), showing a preferential neutralisation with this element. However, the low slope of this relation ( $a = 0.15$ ) and the acid pH distribution (Fig. 4) may indicate that it is poorly effective. These observations are reliable with the results of Plaisance et al. (1997) who show the existence of  
245 important sources of  $\text{NO}_2$  and  $\text{SO}_2$  in the industrial and urban regions of northern and central Europe, showing the importance of anthropogenic emissions to the high acid events that occur in France.

The chemistry of rainfall events from sector 3 is dominated by non sea-salt calcium (28.4  
250  $\mu\text{eq/l}$ ) and bicarbonate (26.6  $\mu\text{eq/l}$ ), which could be attributed to dissolution of calcareous dust in rain originated from Mediterranean region (Plaisance et al., 1996; Avila et al., 1997; Avila et al., 1998). The weighted mean concentration of ammonium, non-sea-salt sulphates and nitrates are relatively low: 17.2  $\mu\text{eq/l}$ , 16.0  $\mu\text{eq/l}$ , 16.9  $\mu\text{eq/l}$ , respectively. Distribution of pH is similar to the one of sector 1. For these two groups, the neutralisation effect maintains  
255 pH close to the neutral threshold. Fig. 5 indicates that neutralisation is mainly due to  $\text{NH}_4^+$  ( $R = 0.86$ ) for sector 3. Weak concentrations in anthropogenic components are not in agreement with the observation of Ezcurra et al. (1988) and Diaz-Caneja et al. (1989). These authors quoted that polluted rains for northern Spain were mostly associated to industrial sources in Southern France and Northern Spain. This difference could be explained by the presence of  
260 the Cevennes massif located between Mediterranean area and sampling site and that plays the role of an orographic barrier. Such hypothesis is consistent with the low concentration in marine species. Indeed, if air-masses from the South are partially stopped and precipitated, it would mean that most of clouds precipitating at Opme were supplied by closer sources of

aerosols mainly emitting calcareous material. This assumption is reliable with the proportion  
265 between  $\text{nssCa}^{2+}$  and the acidic species ( $a = 0.93$ ).

Sector 4 is represented by rainfalls events produced by local influences. This group present  
high weighted mean value in non-sea-salt calcium (36.8  $\mu\text{eq/l}$ ), bicarbonates (32.2  $\mu\text{eq/l}$ ),  
non-sea-salt sulphates (31.0  $\mu\text{eq/l}$ ), ammonium (31.8  $\mu\text{eq/l}$ ) and nitrates (29.7  $\mu\text{eq/l}$ ). This  
270 tendency underlines the influence of the local environment, as Opme is located near  
Clermont-Ferrand urban district and Limagne basin that could also provide anthropogenic and  
terrestrial material from cultivated lands. Sea-salt chloride and sodium present the lowest  
mean concentration (8.0 and 8.6  $\mu\text{eq/l}$ , respectively) of the four sectors, due to the inland  
origin of air-masses. Correlation coefficients between  $[\text{NO}_3^- + \text{nssSO}_4^{2-}] / \text{nssCa}^{2+}$  ( $R = 0.74$ )  
275 and  $[\text{NO}_3^- + \text{nssSO}_4^{2-}] / \text{NH}_4^+$  ( $R = 0.77$ ) show that neutralisation is non-specific for sector 4  
and explain that more than 50% of the events of sector 4 present a  $\text{pH} > 5.5$ .

Neutralisation processes during air masses travel are then very efficient, especially for sectors  
1, 3 and 4 due to intervention of  $\text{nssCa}^{2+}$  and  $\text{NH}_4^+$  that act together. Concerning the sector 2,  
280 the weak neutralisation of acidity could be due to the weak proportion of  $\text{NH}_4^+$  content  
according with acidic species. Moreover, the dissolution of ammonium could lead to enhance  
the solubility of  $\text{SO}_2$ . This could constitute a negative feedback on the neutralisation process  
(Finlayson-Pitts et Pitts, 1986).  $\text{Ca}^{2+}$  appears to be the dominant ion of all the sectors. Then,  
among specific characteristics of each sectors, observation show the importance of long-range  
285 transport on rainfall chemistry at Opme météo-station.

Influence of local scavenging: washout

Sequential sampling enables the investigation of the variability of the chemical composition,  
290 and thus the evaluation of scavenging processes (rainout or washout) during rain events  
(Meszaros, 1981; Seymour and Stout, 1983; Durana et al., 1992; Lim et al., 1991; Pelicho et  
al., 2006; Germer et al., 2007; Celle-Jeanton et al., 2009). For this purpose, 29 rain events  
were sampled during the period Dec 2005- May 2006 using a sequential rain collector.

Our purpose is to follow the evolution of rainfall chemistry during rain event and to determine  
295 the local contribution during washout. Taking into account the whole data set, Fig. 6 shows  
that the decrease of concentration concerns all the measured species, except  $\text{H}_3\text{O}^+$ . An average  
of 53% of the total mineralization is removed in the first three fractions (precipitation height =  
2.6 mm), this value varies from 28% to 76%. This indicates that the below cloud scavenging  
generally dominates the removal of atmospheric components in the region. Then a further  
300 gradual drop to approximately constant values is observed and represents the in-cloud  
processes contribution to rainwater chemistry (Gonçalves et al., 2002).

In order to understand associations between chemical species during events and to determine  
the role of meteorological features, four of these 29 rainfalls have been selected according to  
305 their similar rainwater amount (between 7.1 and 9.7 mm) and air mass back trajectory: rainfall  
of 23-24/03/2006 belonging to sector 1, 11/04/2006 to sector 2, 06/05/2006 to sector 3,  
07/05/2006 to local sector. Table 4 presents statistics of rainfall amount, electrical  
conductivity, pH, direction (in degree,  $0^\circ$  is the North;  $180^\circ$  the South) and speed of wind for  
each selected event. Fig. 7, 8, 9 and 10 show the evolution of rainfall height and chemical  
310 content during the 4 selected rain events.

The first fraction of the 23-24/03/06 event, coming from sector 1 (Fig. 7), is characterised by  
the dominance of the  $\text{NO}_3^-$ ,  $\text{nssSO}_4^{2-}$  and  $\text{NH}_4^+$  (49%) whereas  $\text{nssCa}^{2+}$ ,  $\text{nssMg}^{2+}$  and  $\text{HCO}_3^-$

account for 29% and  $\text{Na}^+$  associated with  $\text{ssCl}^-$  and sea-salt fractions of  $\text{Cl}^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  
315  $\text{K}^+$  for 22%. Last fraction of sector 1 is dominated by the terrestrial source (45%). This could  
be due to local anthropogenic influences. The relatively low acidity measured in all the sub-  
samples could be explained by neutralisation process by terrestrial species. Marine  
contribution is almost constant; a small decrease could be quoted from the beginning of the  
rain-event (22%) to the end (19%).

320  
First and last fractions of the 11/04/2006 event, coming from sector 2 (Fig. 8), show a  
predominance of anthropogenic components. In between, from fractions 4 to 6, contribution  
of anthropogenic source decreases whereas marine and terrestrial source supplies are  
relatively constant. At the end of the event, an increase in  $\text{NO}_3^-$ ,  $\text{NH}_4^+$ ,  $\text{nssSO}_4^{2-}$  and  $\text{H}_3\text{O}^+$  and  
325 the absence of decrease of the rainfall intensity may indicate a local recharge of  
anthropogenic and terrestrial species. The significant value of the wind speed during the event  
(5.5 m/s) coming from the North would indicate the influence of Clermont-Ferrand urban  
district. Moreover a rapid raise of the acidity is observed at the end of the event and attests  
that neutralisation of acidity can be observed preferentially at the beginning of a precipitation  
330 event due to carbonate dissolution.

The first fraction of the 06/05/2006 event, coming from sector 3 (Fig. 9), is marked by high  
 $\text{NO}_3^-$ ,  $\text{nssSO}_4^{2-}$ ,  $\text{nssCa}^{2+}$  and  $\text{HCO}_3^-$  content. The first sub-sample is the result of little rainfall  
episodes occurring before the beginning of the major part of precipitation. The chemical  
335 content decreases from fraction 1 to 5. An interruption of the rain, between 13h15 and 13h45,  
allows a reloading of the below-cloud atmosphere that affects more peculiarly anthropogenic  
components. This fact can be related to a possible contribution of local sources of pollutants  
or a reloading of clouds by long range transported components; although the air-mass that

generates the whole event of the 06/05/2006 is coming from the South, for the period  
340 comprised from 13h and 14h30, the wind comes from the North.

The 07/05/2006 event due to local weather pattern (Fig. 10) shows high concentrations in  $\text{HCO}_3^-$  and  $\text{NH}_4^+$ . From the beginning to the end of the rainfall event, concentrations of all components decrease, except  $[\text{H}_3\text{O}^+]$ . This indicates an efficiency of the below cloud  
345 scavenging process. Increasing acidity can be due to the removal of carbonated species below cloud as it can be seen with the decrease of non sea-salt calcium, bicarbonates and ammonium. Indeed, the concentration of this latter is three times lower at the end of the event than in the first sample.

The slight increase of the concentrations in  $\text{NO}_3^-$ ,  $\text{NH}_4^+$ ,  $\text{nssSO}_4^{2-}$  for the 7<sup>th</sup> and the 8<sup>th</sup>  
350 samples corresponds to lower rainfall intensity (e.g. coming from less diluted parts of cloud) whereas concentrations in other components remain constant. This implies a great contribution of anthropogenic species as long-range aerosol loading i.e. condensation nuclei.

Changes in solute concentrations over time within events are likely to be a function of the  
355 relative fraction of fine to coarse aerosols (Germer et al., 2007). The fine aerosol fraction ( $< 5 \mu\text{m}$ ) is predominant within clouds and responsible for rainout (Lim et al., 1991), while the coarse fraction of aerosols, which is subject to gravitational deposition, is more important in washout (Seinfeld and Pandis, 1998; Saha and Moorthy, 2004). Indeed, the collection efficiency varies mainly with the size of the aerosol particles (Chate and Kamra, 1997; Chate  
360 et al., 2003). Gaseous species such as  $\text{NO}_x$ ,  $\text{SO}_x$  and  $\text{NH}_3$  are then predominantly removed during in-cloud processes, whereas coarse mode (terrestrial species) is efficiently deposited during washout. Wet deposition of  $\text{H}_3\text{O}^+$  increases during the event, the highest part being deposited in the last precipitation sample fractions. This increasing acidification may be due

to the high neutralisation effect at the beginning of the event. This effect can happen when  
365 raindrops impact and dissolve atmospheric particles containing alkaline substances, especially  
carbonates.

Neutralising species can be also associated with  $\text{NO}_3^-$ ,  $\text{SO}_4^{2-}$  and  $\text{NH}_4^+$ , forming  $\text{NH}_4\text{NO}_3$ ,  
( $\text{NH}_4$ ) $_2$  $\text{SO}_4$ ,  $\text{Ca}(\text{NO}_3)_2$ ,  $\text{CaSO}_4$ , as can be seen from the important flux of anthropogenic  
370 species such as non sea-salt calcium in the first sample. It would be consistent with the fact  
that anthropogenic species remain an important source and that acidity flux is low at the  
beginning of the event. This fact is linked to the acidity evolution during the four events. The  
major  $\text{H}_3\text{O}^+$  flux occurs during the last fractions of the rainfall event. The rural environment  
of Opme consists of agriculture land and calcareous Oligocene fields and is then an important  
375 source of neutralising species (calcium carbonate). This local characteristic may justify the  
dominance of calcium in the composition of rain coming from the four air-masses patterns.  
The proximity of Clermont-Ferrand agglomeration can also induce anthropogenic feature of  
the rainwater at the beginning or during the event. These facts highlight the importance of  
local meteorological conditions (speed and direction of wind, humidity of ground) to  
380 understand chemical signature of a single rain event.

#### 4. Conclusion

Based on 217 rain samples, this study primarily shows that rainfall chemistry at Opme meteo-  
385 station is dominated by terrestrial elements with a quite neutral pH. The statistical approach  
through a factor analysis, revealed that neutralisation is the main process affecting the  
rainwater chemistry, in agreement with pH measurements.



Taking into account the main air-masses pattern at Opme, this work reveals differences in solute concentrations according to geographical long-range origin. Indeed, whereas sector 1 (Atlantic) and sector 3 (Mediterranean) events are the most diluted, the concentrations increase markedly in the events coming from the sector 4 (local) and the sector 2 (Northern and Eastern Europe). This is probably due to the high flux of aerosols in this industrial area of Europe. In particular it was shown that nitrate and acidity are preferentially carried by these rain events. The highest concentrations of marine species ( $\text{Na}^+$ ,  $\text{ssCl}^-$ ) and in  $\text{nssCa}^{2+}$  are also measured for sector 2. Concerning the neutralisation, it was demonstrated that this process is unspecific (involving both  $\text{nssCa}^{2+}$  and  $\text{NH}_4^+$ ) for sector 1, 3 and 4, but it concerns only  $\text{NH}_4^+$  for sector 2. This characteristic can restrain the neutralisation process and account for the acid pH distribution of this last sector.

Considering the chemical variability during four single events, it appears that the major part of chemical content is scavenged during the three first fractions of the rainfall. This highlights the great influence of washout on the final composition of the rainwater. It especially contributes to the deposition of large terrestrial aerosols, which neutralise acidity of the first fractions of rainwater and anthropogenic species coming from Clermont-Ferrand urban district. Then, at an infra-event scale, local meteorological phenomena could strongly alter the long-range signatures and account for leading factors of chemical variability of the meteoric water.

### 3.1.1 Acknowledgments

The authors thank the PREVOIR project financed by the Auvergne region for its material support and gratefully acknowledge the NOAA Air Resources Laboratory (ARL) for the provision of the HYSPLIT transport and dispersion model and/or READY website (<http://www.arl.noaa.gov/ready.html>) used in this publication.

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Tables

	E.C (μS/cm)	pH	HC0 <sub>3</sub> <sup>-</sup>	Cl <sup>-</sup>	NO <sub>3</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	Na <sup>+</sup>	NH <sub>4</sub> <sup>+</sup>	K <sup>+</sup>	Mg <sup>2+</sup>	Ca <sup>2+</sup>
Arithmetic mean	17.5	5.4	32.2	23.0	39.2	28.8	0.7	23.9	28.5	5.5	7.8	39.9
Arithmetic s.d.	14.3	0.7	38.1	29.5	47.4	27.2	3.9	32.7	19.5	6.7	8.2	60.0
Minimum	2.0	3.4	0.0	0.8	0.0	0.0	0.0	0.6	0.0	0.1	0.3	2.0
Maximum	96.0	7.1	200.0	211.4	324.9	134.8	37.4	250.5	92.5	41.2	54.0	413.3
V-W mean	12.6	5.1	27.5	11.5	25.8	24.6	0.3	12.0	24.4	3.1	5.2	33.8
V-W s.d	9.2	0.7	30.4	17.3	30.5	20.8	2.6	19.6	14.5	4.1	5.4	46.0

Table 1: Major ionic constituents (μeq·l<sup>-1</sup>) of the 217 sampled rains during the study period (from Nov-2005 to Nov-2007)

	<b>F1</b>	<b>F2</b>	<b>F3</b>
H <sub>3</sub> O <sup>+</sup>	0.111	0.041	-0.420
HCO <sub>3</sub> <sup>-</sup>	0.326	0.044	<b>0.913</b>
Cl <sup>-</sup>	0.081	<b>0.997</b>	0.005
NO <sub>3</sub> <sup>-</sup>	<b>0.894</b>	0.207	-0.095
SO <sub>4</sub> <sup>2-</sup>	<b>0.874</b>	0.257	0.125
PO <sub>4</sub> <sup>3-</sup>	0.338	-0.042	0.184
Na <sup>+</sup>	0.124	<b>0.920</b>	-0.088
NH <sub>4</sub> <sup>+</sup>	<b>0.764</b>	0.146	0.035
K <sup>+</sup>	0.341	0.260	0.300
Mg <sup>2+</sup>	0.480	<b>0.711</b>	0.305
Ca <sup>2+</sup>	<b>0.532</b>	0.184	<b>0.606</b>
% of total variance	41.7	35.6	22.4
	Neut.	Marine	Terrestrial

**Table 2: Varimax rotated factor loadings, total variance and the determination of the main different sources**

	Sector 1		Sector 2		Sector 3		Sector 4	
<b>Nb of events</b>	75		49		51		42	
<b>Height</b>	309		179		303		248	
	V-W-Mean	V-W-Std dev	V-W-Mean	V-W-Std dev	V-W-Mean	V-W-Std dev	V-W-Mean	V-W-Std dev
<b>E.C.</b>	10.2	8.0	19.5	12.7	9.6	6.4	14.3	7.0
<b>pH</b>	5.1	0.7	4.8	0.7	5.1	0.7	5.3	0.6
<b>HCO<sub>3</sub><sup>-</sup></b>	24.1	28.9	28.3	31.5	26.6	37.2	32.2	21.3
<b>NO<sub>3</sub><sup>-</sup></b>	16.0	15.2	52.3	53.8	16.9	13.6	29.7	22.4
<b>PO<sub>4</sub><sup>3-</sup></b>	0.0	0.2	1.4	5.6	0.0	0.3	0.2	2.2
<b>ssCl<sup>-</sup></b>	14.7	23.6	13.8	30.1	8.6	11.7	8.0	6.6
<b>nssCl<sup>-</sup></b>	0.3	11.4	0.7	18.0	0.3	10.6	0.1	11.6
<b>ssSO<sub>4</sub><sup>2-</sup></b>	1.7	23.3	1.9	20.5	1.1	10.8	1.0	6.2
<b>nssSO<sub>4</sub><sup>2-</sup></b>	16.8	1.2	35.6	3.9	16.0	1.5	31.0	0.2
<b>Na<sup>+</sup></b>	14.9	2.7	15.6	3.6	9.7	1.4	8.6	0.8
<b>NH<sub>4</sub><sup>+</sup></b>	20.0	18.5	33.3	24.1	17.2	14.0	31.8	18.3
<b>ssK<sup>+</sup></b>	0.3	0.5	0.3	0.6	0.2	0.2	0.2	0.1
<b>nssK<sup>+</sup></b>	2.2	3.0	4.7	6.6	2.2	2.8	3.1	2.9
<b>ssMg<sup>2+</sup></b>	3.0	4.5	2.8	3.2	2.0	2.4	1.9	1.5
<b>nssMg<sup>2+</sup></b>	2.0	2.8	4.7	5.4	2.1	3.2	3.2	2.7
<b>ssCa<sup>2+</sup></b>	0.7	1.0	0.7	1.3	0.4	0.5	0.4	0.3
<b>nssCa<sup>2+</sup></b>	23.9	44.6	53.2	66.0	28.4	40.4	36.8	27.8

**Table 3: Weighted mean value ( $\mu\text{eq}\cdot\text{l}^{-1}$ ) of major ionic rain components for the four meteorological sectors (ss = sea salt / nss = non sea-salt)**

Event	Long-range provenance	Duration	Height (mm)	Number of fractions	Mean Wind speed (m/s)	Mean Wind direction (degrees)	Mean E.C. ( $\mu\text{S}/\text{cm}$ )	Mean pH
<b>23-24/03/06</b>	Sector 1	2 h 15	7.3	8	$5.5 \pm 0.7$	$167 \pm 8$	$16.3 \pm 8.0$	$5.9 \pm 0.2$
<b>11/04/2006</b>	Sector 2	4 h 00	7.1	8	$3.3 \pm 1.1$	$269 \pm 60$	$6.5 \pm 2.9$	$5.5 \pm 0.2$
<b>06/05/2006</b>	Sector 3	4 h 30	9.7	10	$2.0 \pm 1.6$	$134 \pm 75$	$18.0 \pm 10.0$	$5.2 \pm 0.4$
<b>07/05/2006</b>	Sector 4	1 h 00	9.3	10	$2.5 \pm 1.6$	$117 \pm 63$	$7.6 \pm 2.9$	$5.4 \pm 0.4$

**Table 4: Mean characteristics of the four fractionated events**

Figures

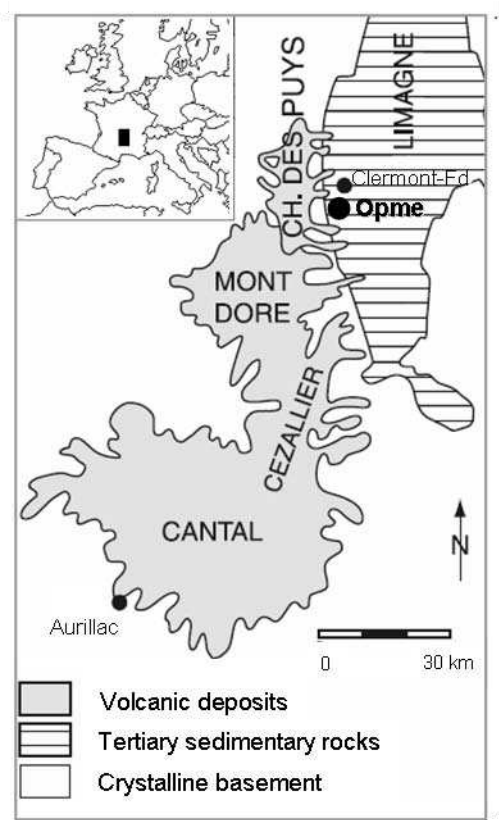


Fig. 1 Location and geological settings of the studied area

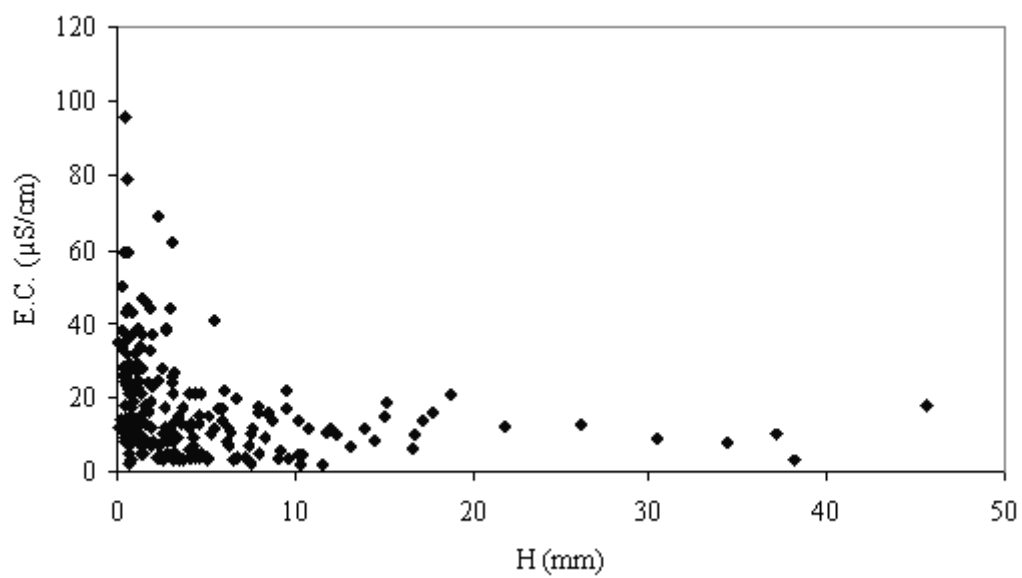
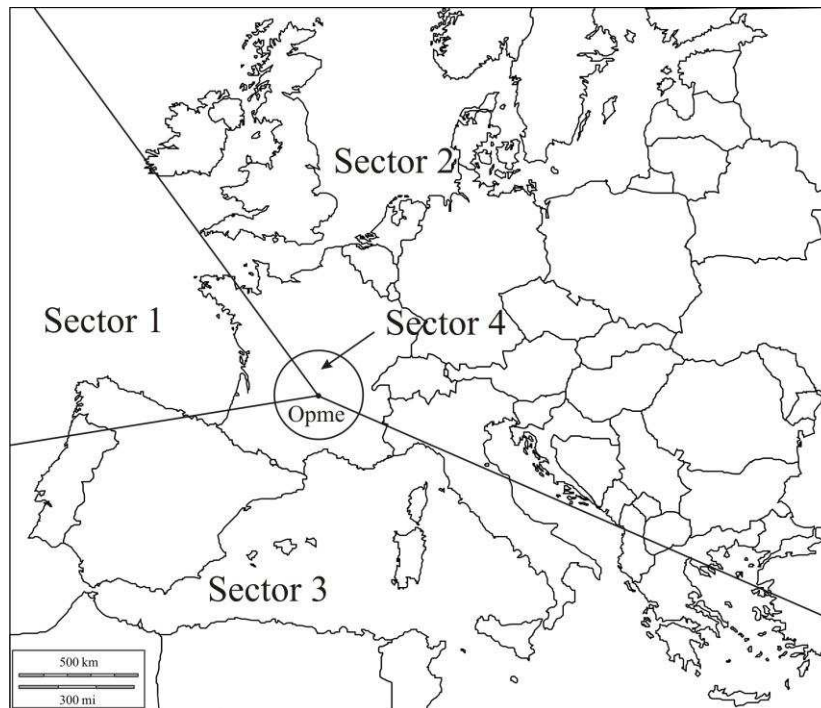
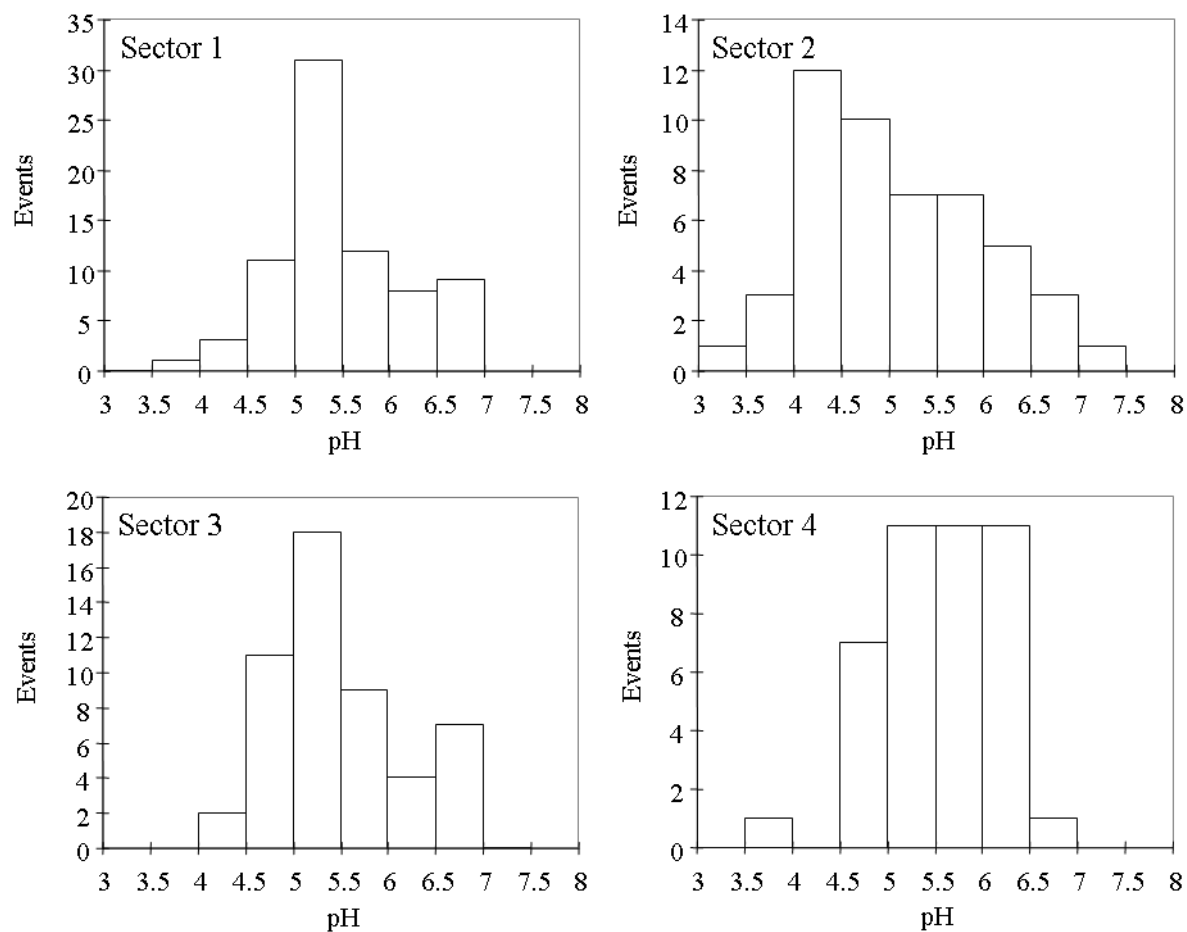


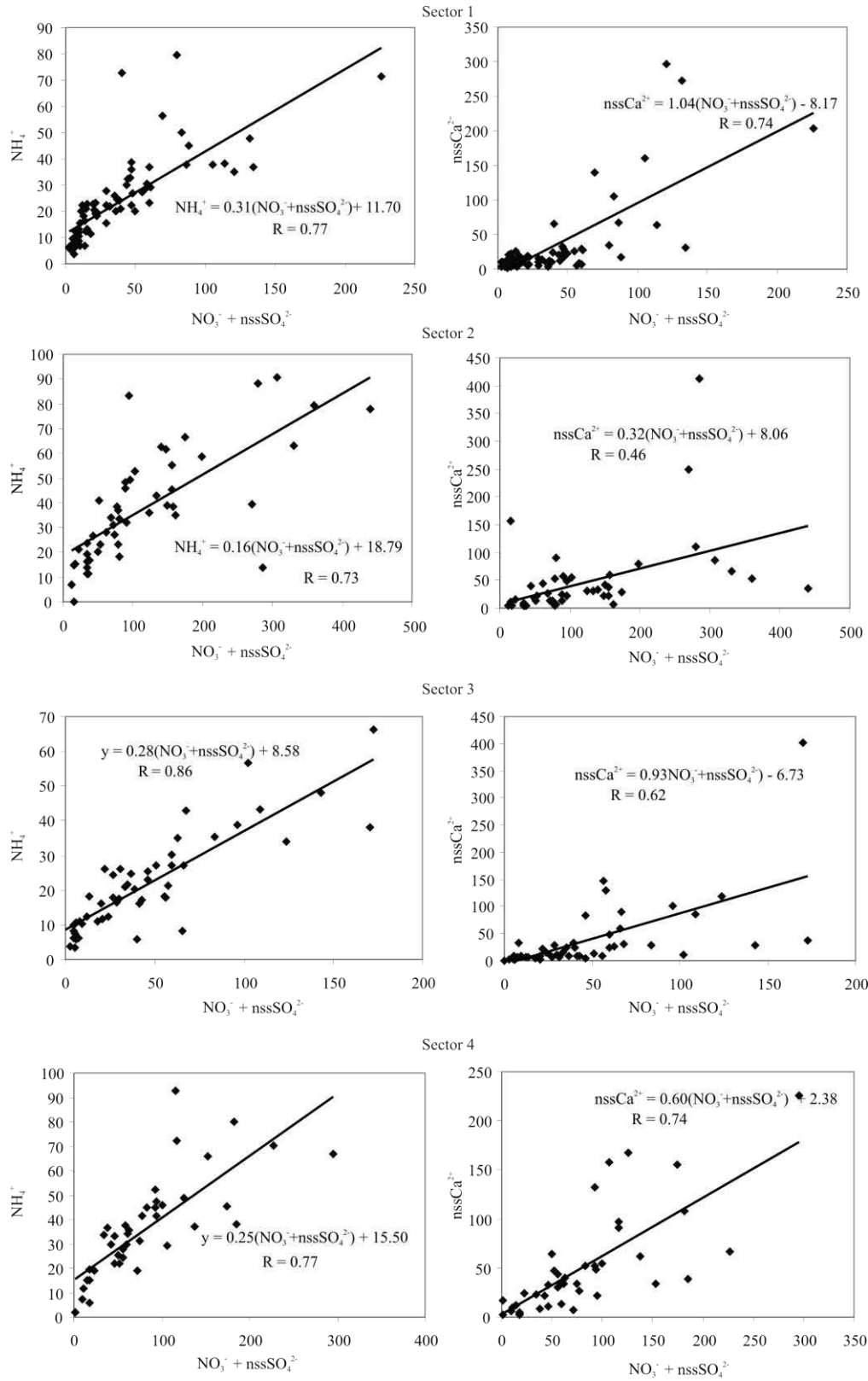
Fig. 2 Relationship between precipitation height (mm) and E.C. (μS/cm) of rainwater



**Fig. 3 Main sectors for air masses trajectories arriving at Opme: (1) West, (2) North and East, (3) South including Iberian and Italian peninsulae and (4) Local**

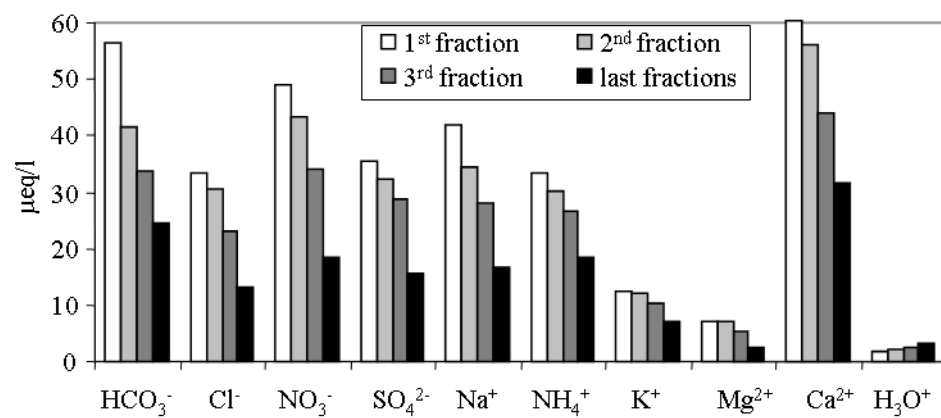


**Fig. 4** pH distribution for the four meteorological sectors

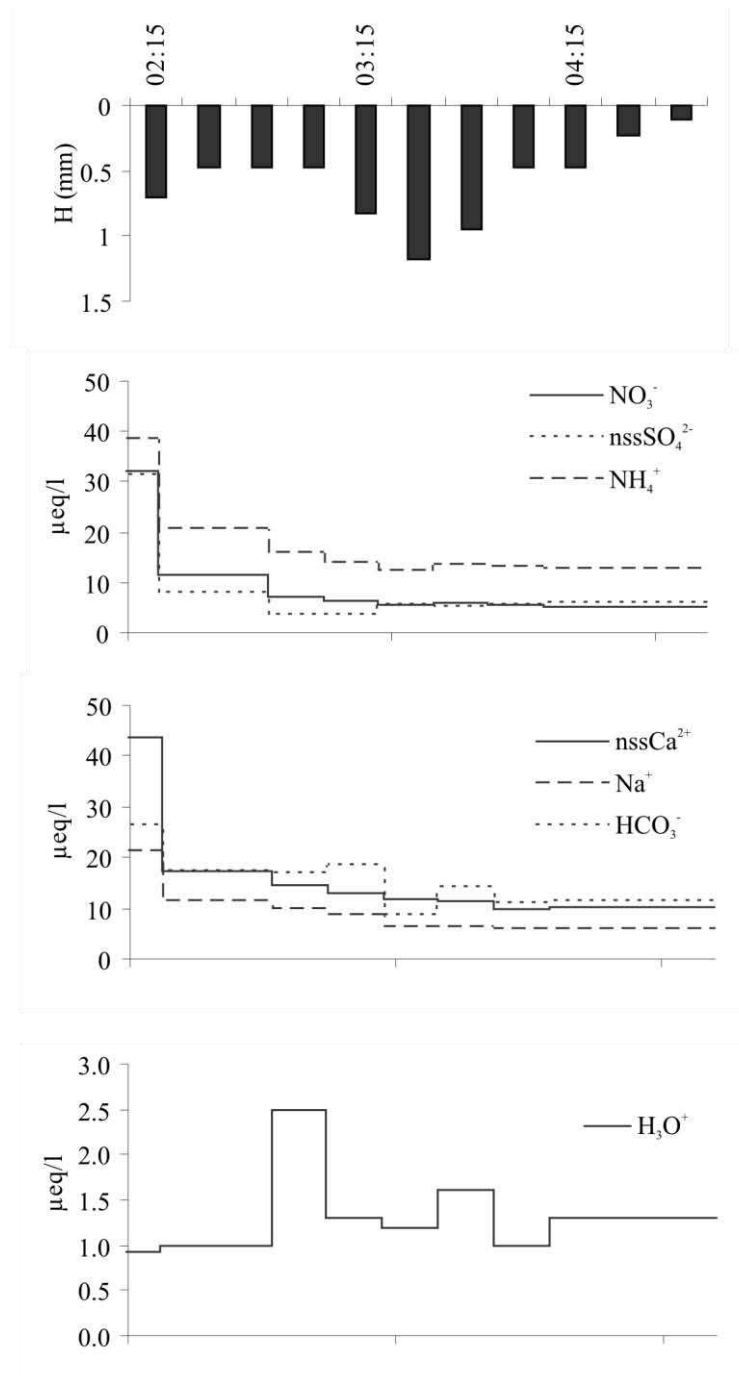


**Fig. 5** Non sea-salt  $\text{Ca}^{2+}$  versus  $[\text{NO}_3^- + \text{nssSO}_4^{2-}]$  and  $\text{NH}_4^+$  versus  $[\text{NO}_3^- + \text{nssSO}_4^{2-}]$  for the four meteorological sectors

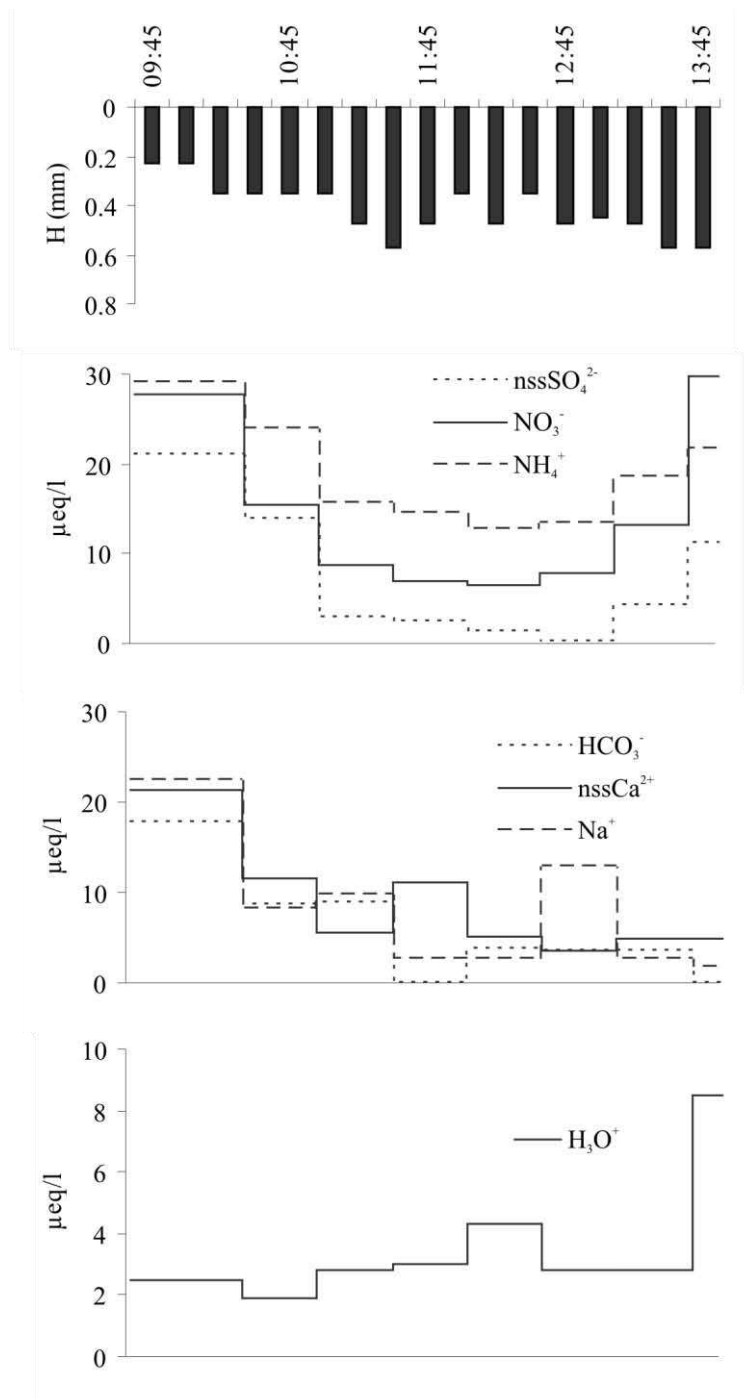




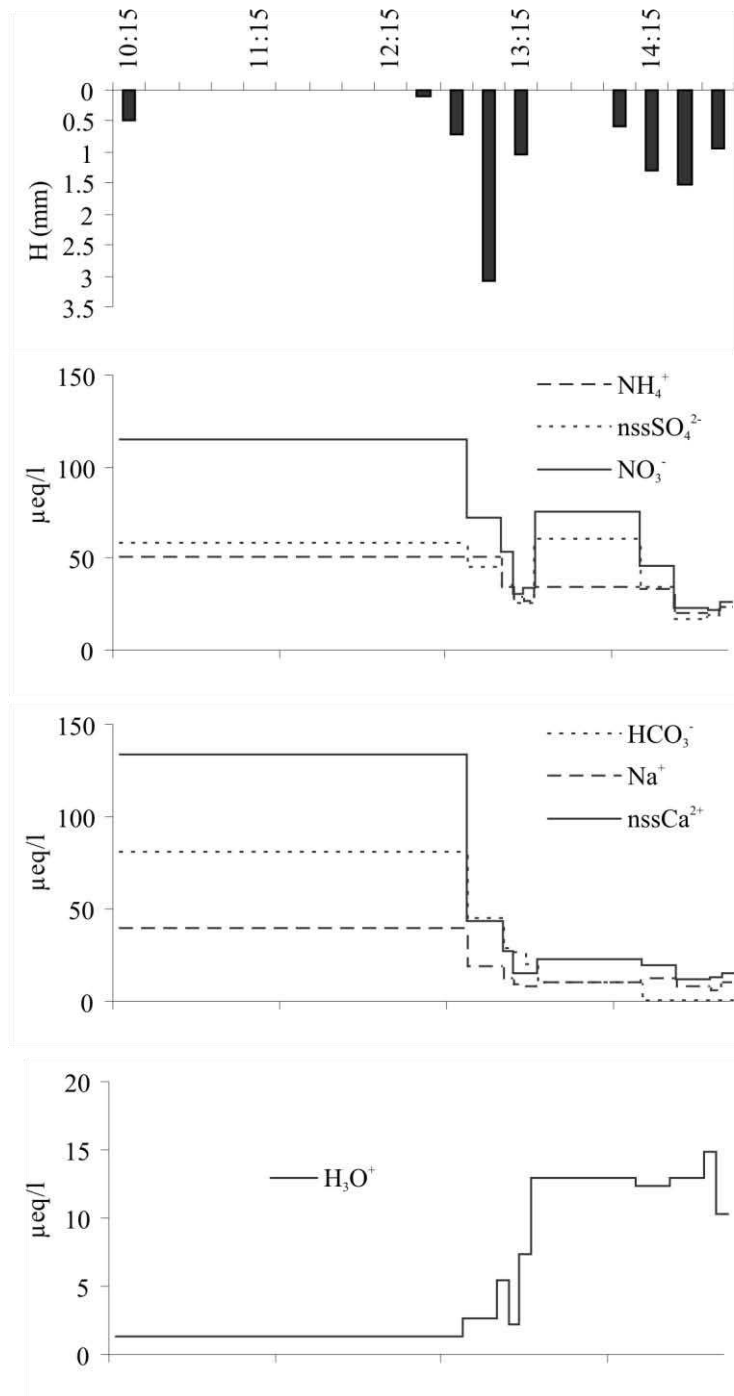
**Fig. 6 Mean chemical content evolution for the 29 fractionated sampled events**



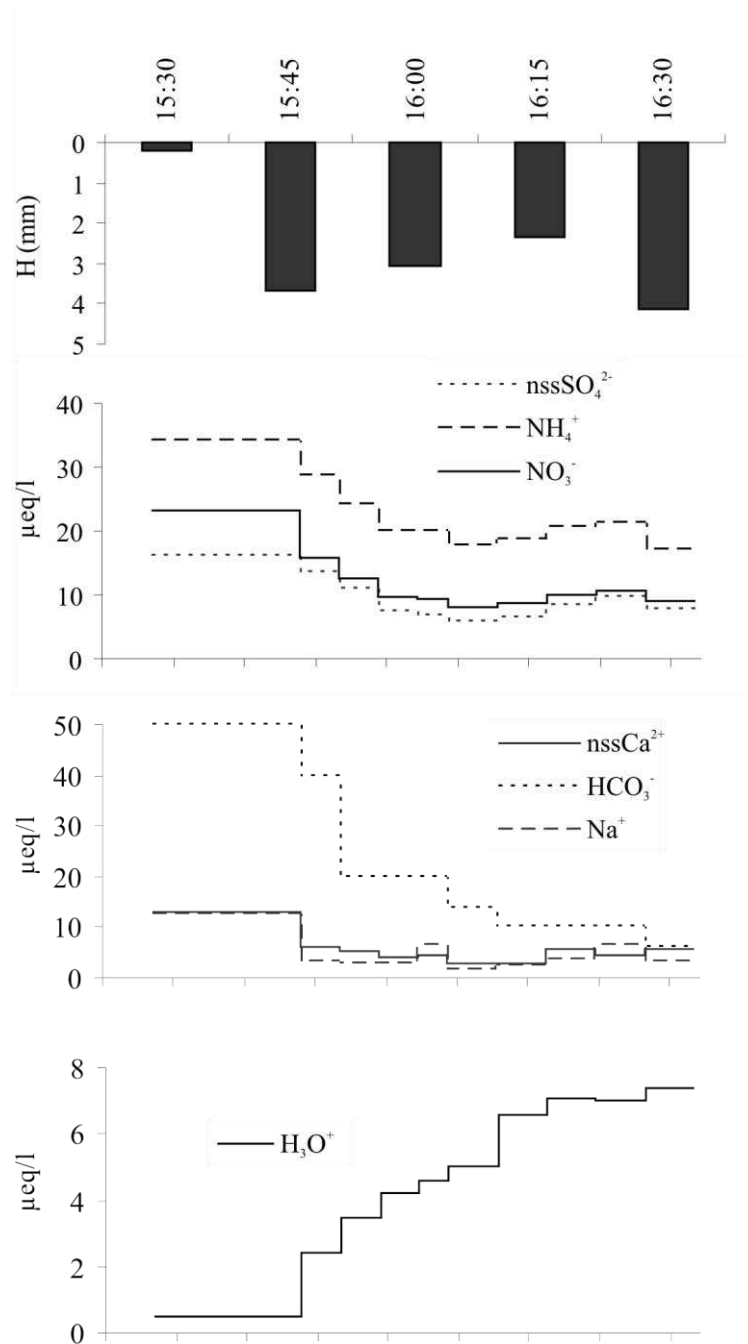
**Fig. 7 Evolution of the chemical content and rainfall height for the 23-24/03/2006 event coming from sector 1**



**Fig. 8 Evolution of the chemical content and rainfall height for the 11/04/2006 event coming from sector 2**



**Fig. 9 Evolution of the chemical content and rainfall height for the 06/05/2006 event coming from sector 3**



**Fig. 10 Evolution of the chemical content and rainfall height for the 07/05/2006 event coming from sector 4**